http://www.eoht.info/page/Thermodynamics+humor

In one episode of the *The Simpsons*, after Lisa constructs a perpetual motion machine whose energy increases with time, Homer scolds her with:

"In this house, we obey the laws of thermodynamics!"

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In the late 1940s, German theoretical physicist Arnold Sommerfeld, having previously written a series of books in physics: mechanics (1943), electrodynamics (1948), optics (1950), etc., was asked why he had never written a book on thermodynamics? The following is his humorous (but fairly accurate) and frequently quoted answer:

"Thermodynamics is a funny subject. The first time you go through it, you don't understand it at all.

The second time you go through it, you think you understand it, except for one or two small points.

The third time you go through it, you know you don't understand it, but by that time you are so used to it, it doesn't bother you anymore."

3 LAWS of THERMODYNAMICS

1st Law: The Energy (of the Universe) is conserved.

2nd Law: The Entropy (disorder) of the *Universe* increases during all natural processes!!!

3rd Law: Entropy of a perfect crystal at zero Kelvin = 0.

(Note the EMPHASIS on 2nd Law)

What is a Law of Nature (or Scientific Law)?

“A scientific Law is a regularity observed in nature and formulated after a very large number of observations”

[by many, many people in many different laboratories at many different times]

“Because they are solidly grounded in experimental observations, scientific laws are sometimes modified after further experience, but they are rarely refuted”

from Page 13 of our Textbook.

There is another law of thermodynamics that is sometimes stated explicitly in textbooks and sometimes not: <u>all of us are constantly doing the experiments that check this Law for its validity:</u>

The Zeroth Law of Thermodynamics:

Heat does not flow spontaneously from a cold object to a warmer object.

This law was invoked in establishing the all-important 2nd Law.

Next: Language and definitions of Thermodynamics

Very important definitions: "system" and "surroundings"

System: whatever you want it to be

Surroundings: everything else

surroundings = (what is ***NOT*** system)

System is anything you can define with a boundary

e.g., a cell, beaker of water, a cat, earth + atmosphere, solar system, etc.

System + Surroundings = "universe"

What do we mean by "universe"?

Even if we enclosed our GALAXY (or even the Solar System) with an insulating, impenetrable shield, we would not notice a difference in our lifetime.

Types of systems (boundaries)

Isolated: Exchanges no energy and no matter
(sealed, insulated, no mechanical connection)
i.e., **a little "universe"**

Closed: Exchanges energy but no matter
(sealed balloon of gas)

Open: Exchanges energy ***and*** matter
(hen's egg, a cell)

More Language (vocabulary) for Thermodynamics

First Law: The ENERGY (U) of the "universe" is constant (conserved)

First Law: $U_{syst} + U_{surr} = U_{universe} = \text{constant}$

notation: $U + U_{surr} = \text{constant}$

(Note: **no subscript** means **system** ;

But, can only know ***CHANGES*** (no such thing as absolute energy)

"Process" means the system changes state: state 1--> state 2
(*different temperature, volume, pressure, chemical nature, etc.*)

state 2: $U_2 + U_{surr2} = U_{univ}$

state 1: $U_1 + U_{surr1} = U_{univ}$

Difference = $(U_2 - U_1) + (U_{surr2} - U_{surr1}) = 0$

$\Delta U + \Delta U_{surr} = 0$ **therefore:** $\Delta U = -\Delta U_{surr}$

$\Delta U = -\Delta U_{surr}$ is a statement of the 1st Law

In 1850: The System was a <u>black box</u> (an engine, a cat, ...)
(They did not know what goes on in the system. Today we do know much more, typically, but must pretend it is a black box when working problems in thermodynamics.)

<u>WE</u> are in the surroundings; we observe energy going in or out of the system.

<u>WE</u> measure two kinds of energy exchange as observed in the SURROUNDINGS.

Heat = q
Work = w

The First Law as the <u>most used</u> equation:

$\Delta U = q + w$ (eq. 2.26, p. 24)

ΔU = heat + work

Energy change of <u>system</u> during process

Energy absorbed by system <u>because of temperature difference</u> between system and surroundings.

Something in the surrroundings <u>*must get colder or hotter,*</u> or else q = 0

Energy absorbed by system because of a <u>*force acting in surroundings*</u>.

Something in surroundings <u>must *move* (or some mechanical happening that could cause something to move</u>, e.g., charging a battery.)

FOR ANY PROCESS:

If q = + we say it is **ENDO**thermic
If q = - we say it is **EXO**thermic

If w = + we say work is done **ON** the system
If q = - we say work is done **BY** the system

In old chemistry books and in most engineering texts one sees:

$$\Delta U = q - w$$

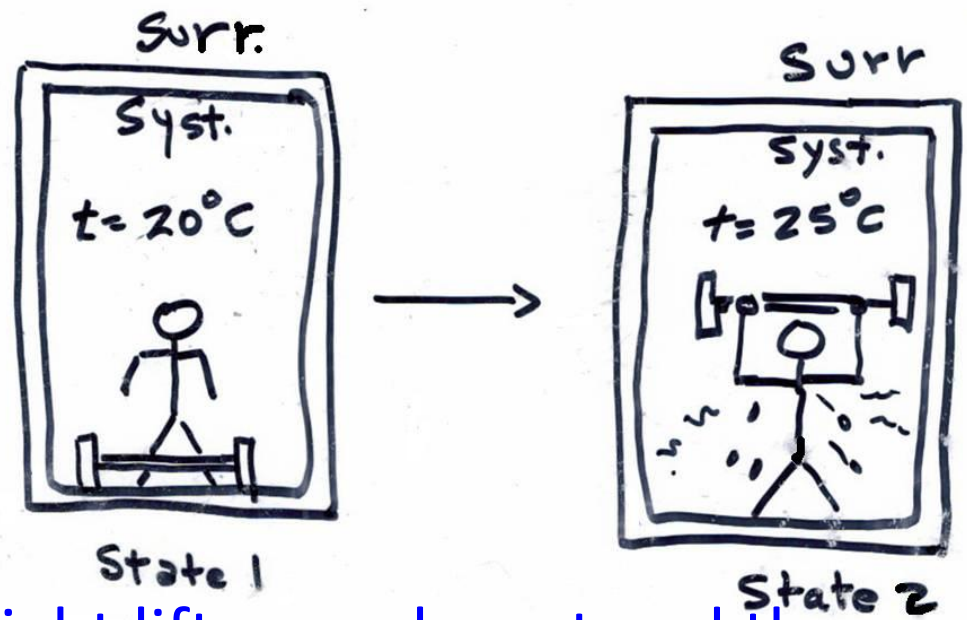
Is this wrong?

What is the definition of w in this case?

w is defined as work obtained FROM the system if we use the above equation

A little quiz

System = **insulated** room with weight lifter



Weight lifter works out and the room heats up from 20° C to 25° C

Are ΔU, q, w positive, zero, or negative?

Answer: all are = 0 because nothing is happening in the surroundings.