Experiment 3: Fluorescence Spectroscopy I (continued)

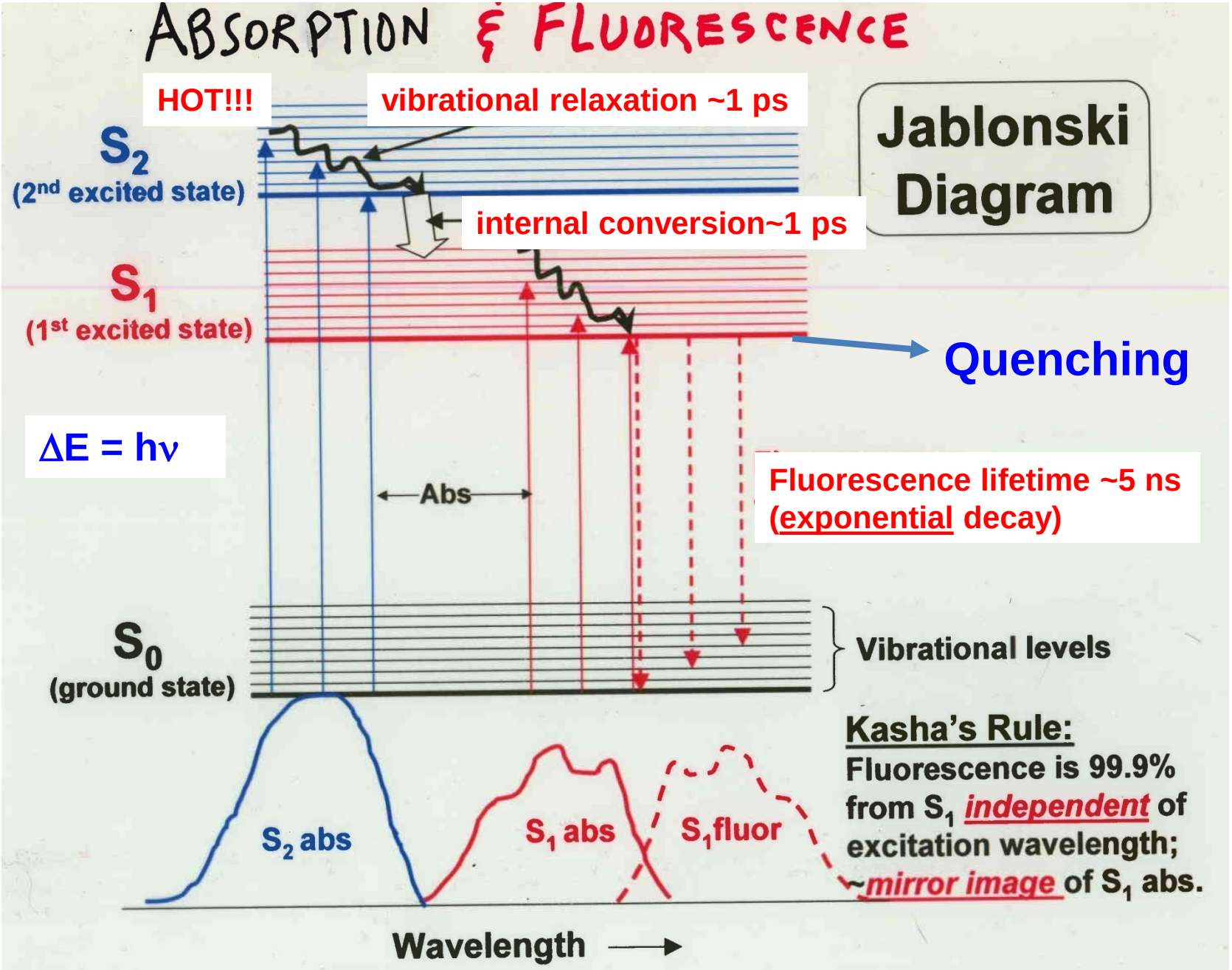
Last week: Part I.A: Introduction to steady state spectra

Today: Part 1.B: Fluorescence Quenching and the Stern-Volmer Relation

Prelab Lecture 7feb18
P. Callis

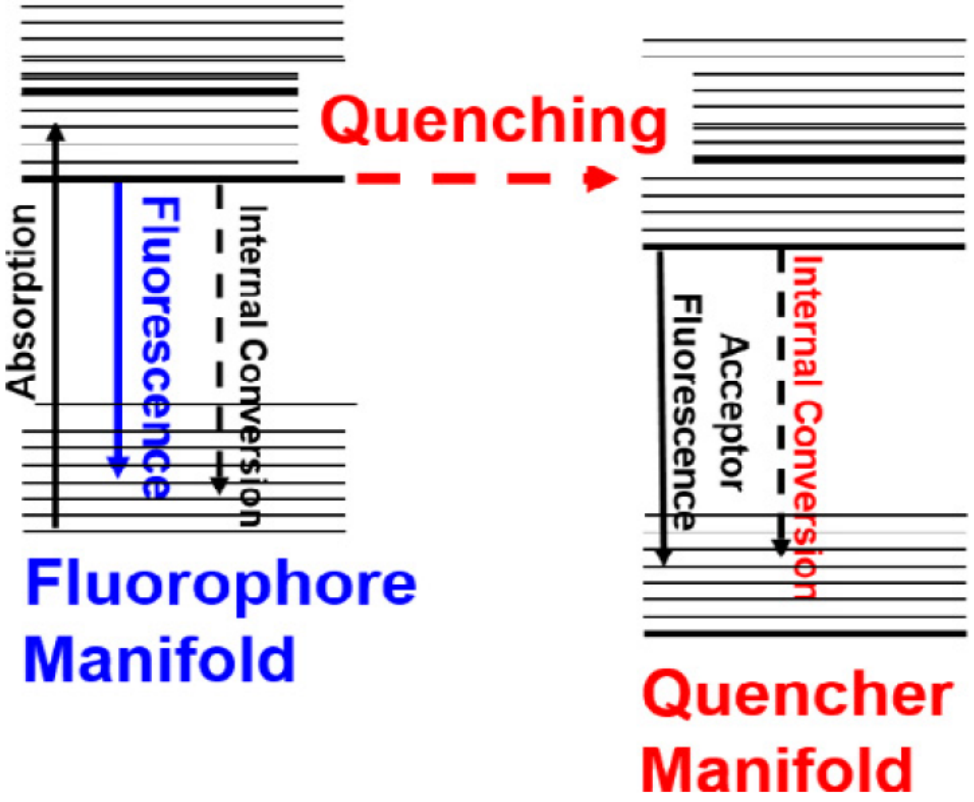
All life appears to be nurtured by the excitation of electrons by light in photosynthesis.

The vision enjoyed by higher life forms begins with the electronic excitation of a conjugated polyene.


ABSORPTION & FLUORESCENCE
HOT!!!
vibrational relaxation ~1 ps
Jablonski Diagram
S2
(2nd excited state)
internal conversion~1 ps
S1
(1st excited state)
Quenching
ΔE = hν
Abs
Fluorescence lifetime ~5 ns
(exponential decay)
S0
(ground state)
Vibrational levels
Kasha's Rule:
Fluorescence is 99.9%
from S1 independent of
excitation wavelength;
~mirror image of S1 abs.
S2 abs
S1 abs
S1fluor
Wavelength

"Quenching" is a generic term usually referring to a process that leads to a decrease in fluorescence....
(***quenching is therefore technically a form of internal conversion***)

Supermolecule Jablonski Diagram (two molecules close together)



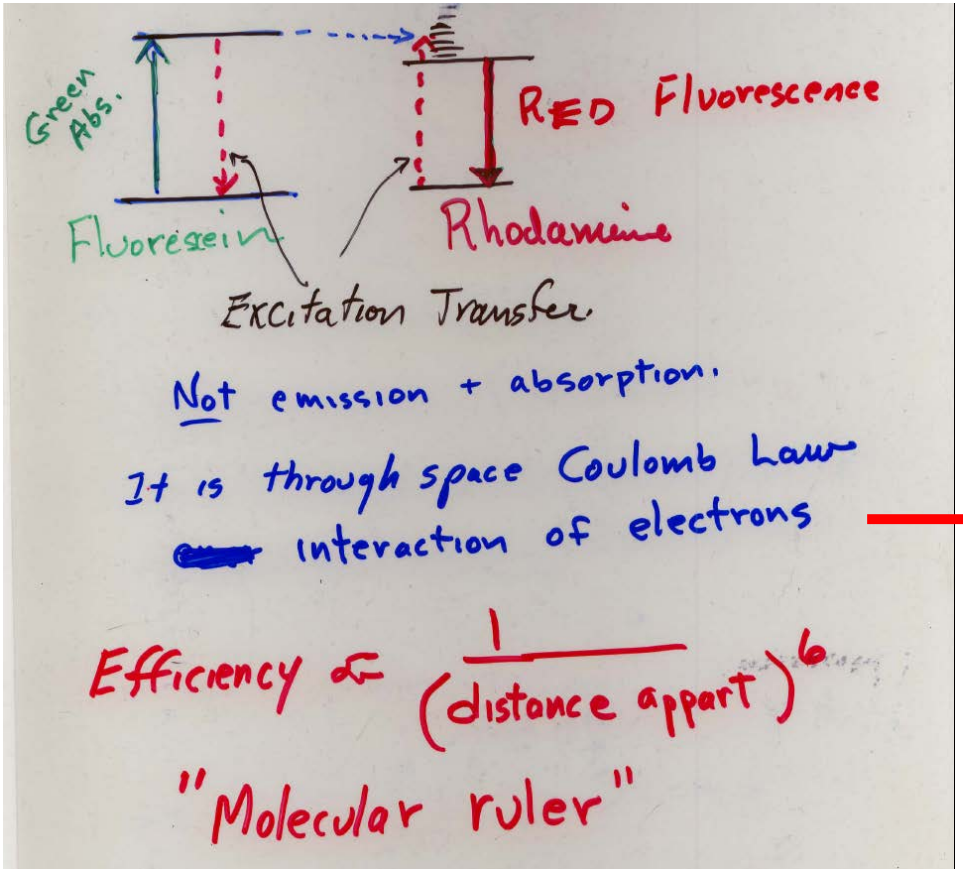
“Quenching and Internal Conversion”

Very important
IN the NATURAL WORLD

Chlorophyll fluoresces if **not** doing photosynthesis;
Using the light to make glucose is **quenching**

Our visual pigments in the retina do fast **internal conversion**, by cis-trans isomerization and **do not fluoresce**

Another kind of quenching:
Forster Resonance Energy Transfer (FRET)



London dispersion force (the very same force that holds liquid nitrogen together.

The FIRST Step of Photosynthesis:
is FRET between "light harvesting" chlorophylls funneling the energy of any absorbed photon to the reaction center chlorophyll.

Binding phenomena and fluorescence quenching. I: Descriptive quantum principles of fluorescence quenching using a supermolecule approach ☆

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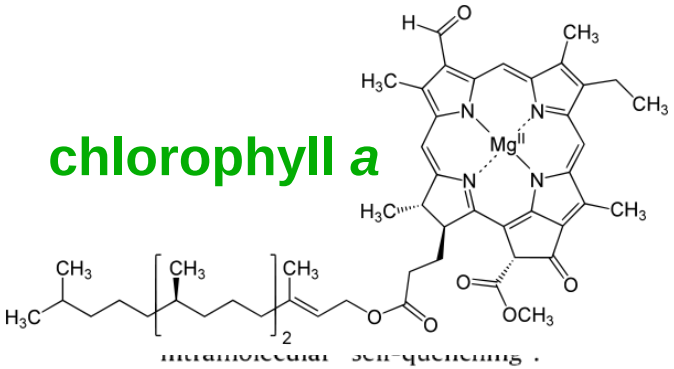
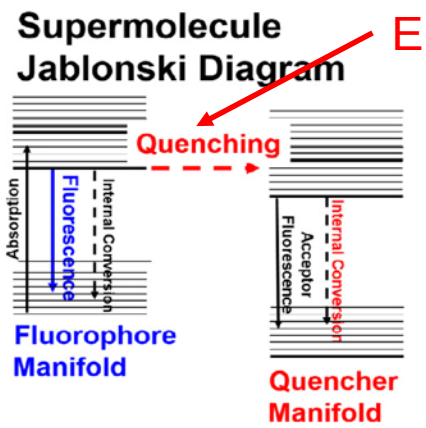
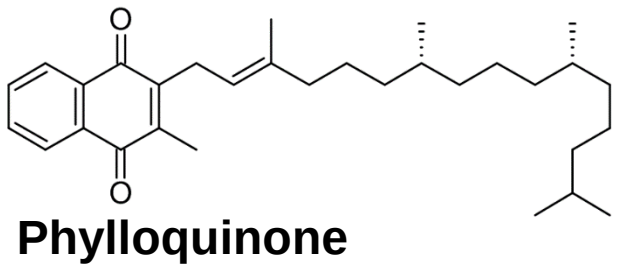
HIGHLIGHTS

GRAPHICAL ABSTRACT

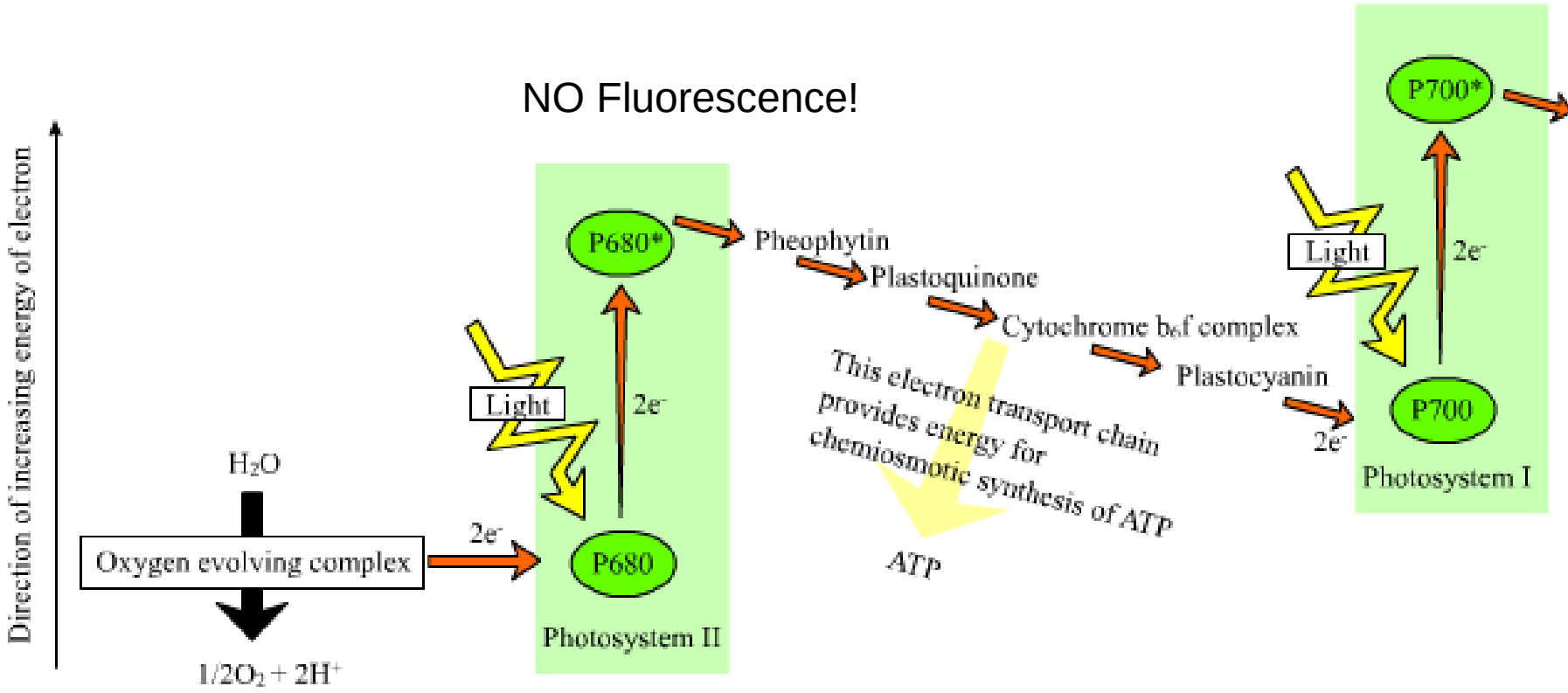






The Second Step of Photosynthesis:
is quenching of the fluorescence of chlorophyll
by fast electron transfer

The Second Step of Photosynthesis: is quenching of the fluorescence of chlorophyll



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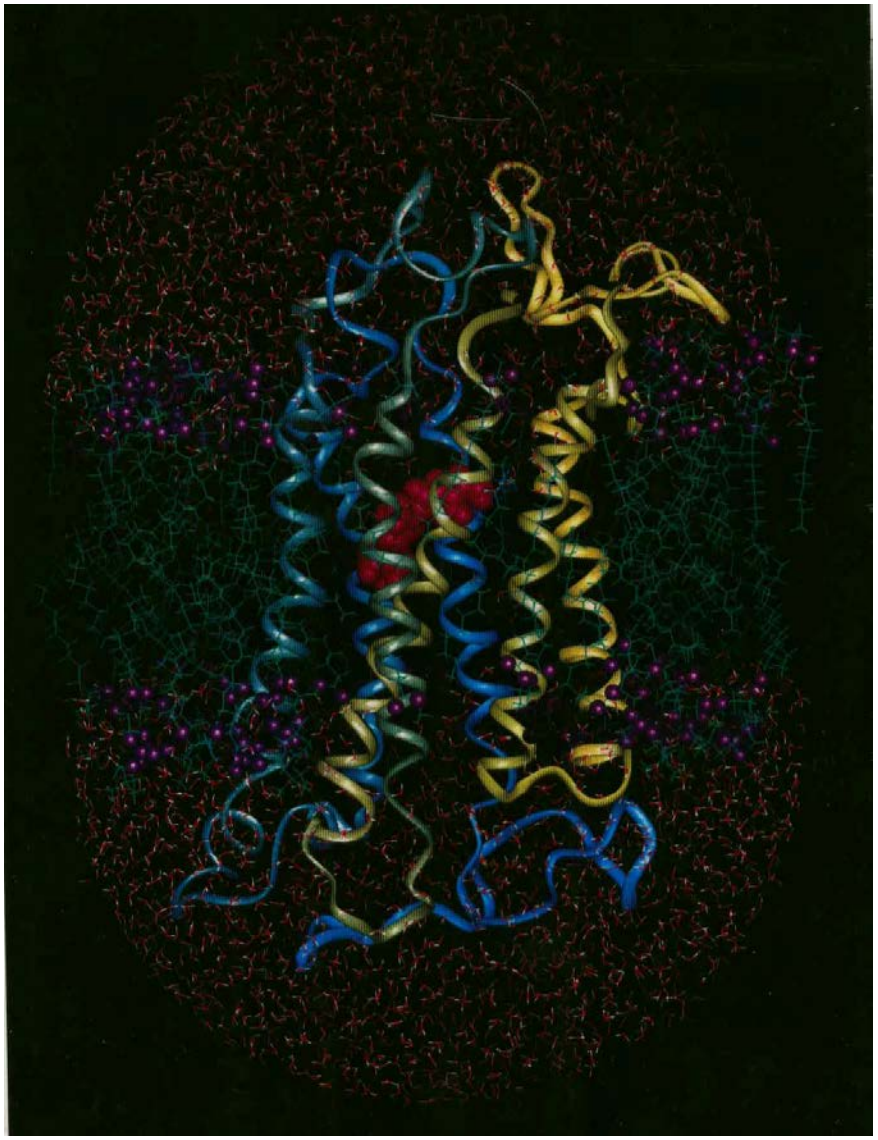

The vision enjoyed by higher life forms begins with the electronic excitation of a conjugated polyene.

This is actually a form of INTERNAL CONVERSION not quenching

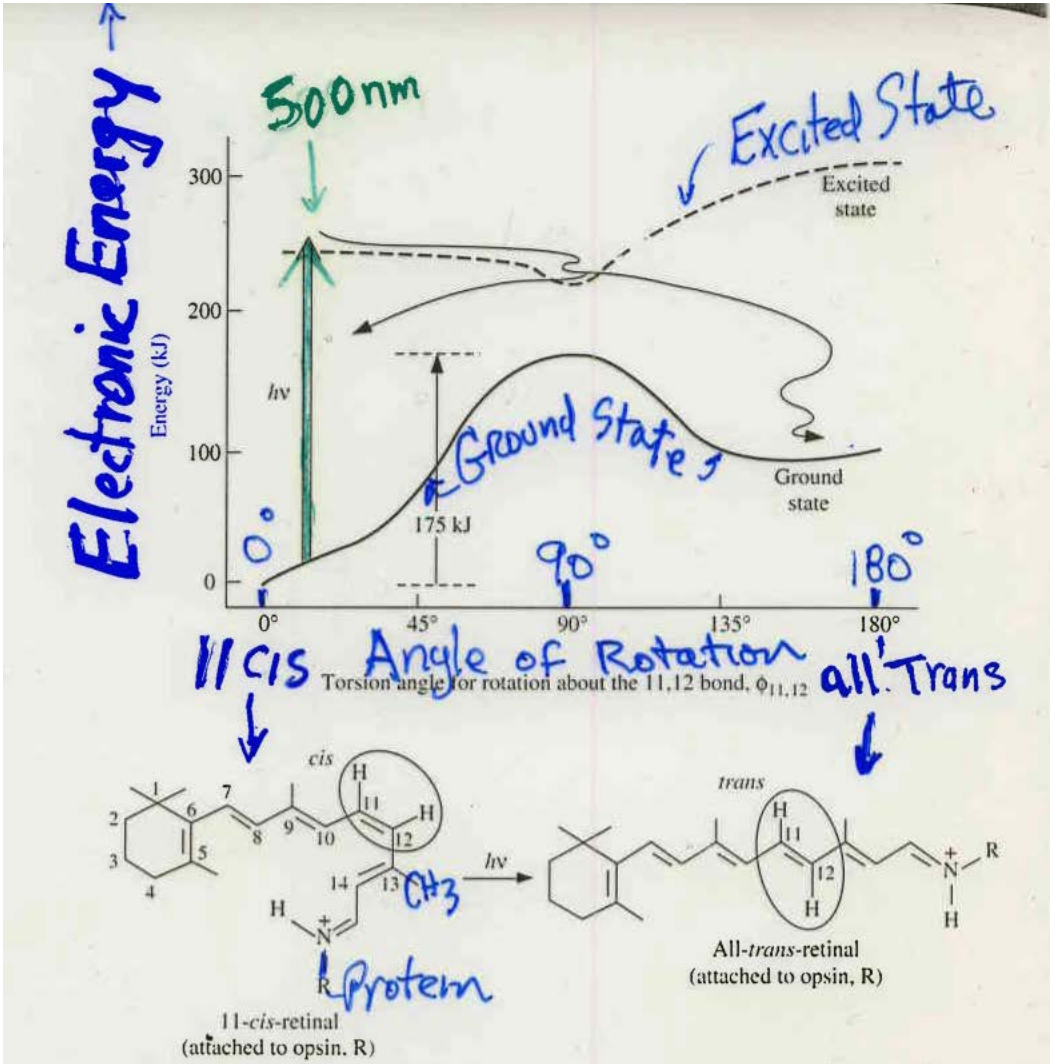

Fig. 9.1 Light-driven isomerization of 11-*cis*-retinal attached to opsin (R) by a Schiff's base linkage. Theoretical calculations of the energy of both the ground state and an excited electronic state suggest a possible path whereby absorption of a photon will lead to conversion from the *cis* to the *trans* configuration about 60% of the time. The angle of rotation about the 11,12-double bond is plotted as the horizontal axis.

Figure 11.3
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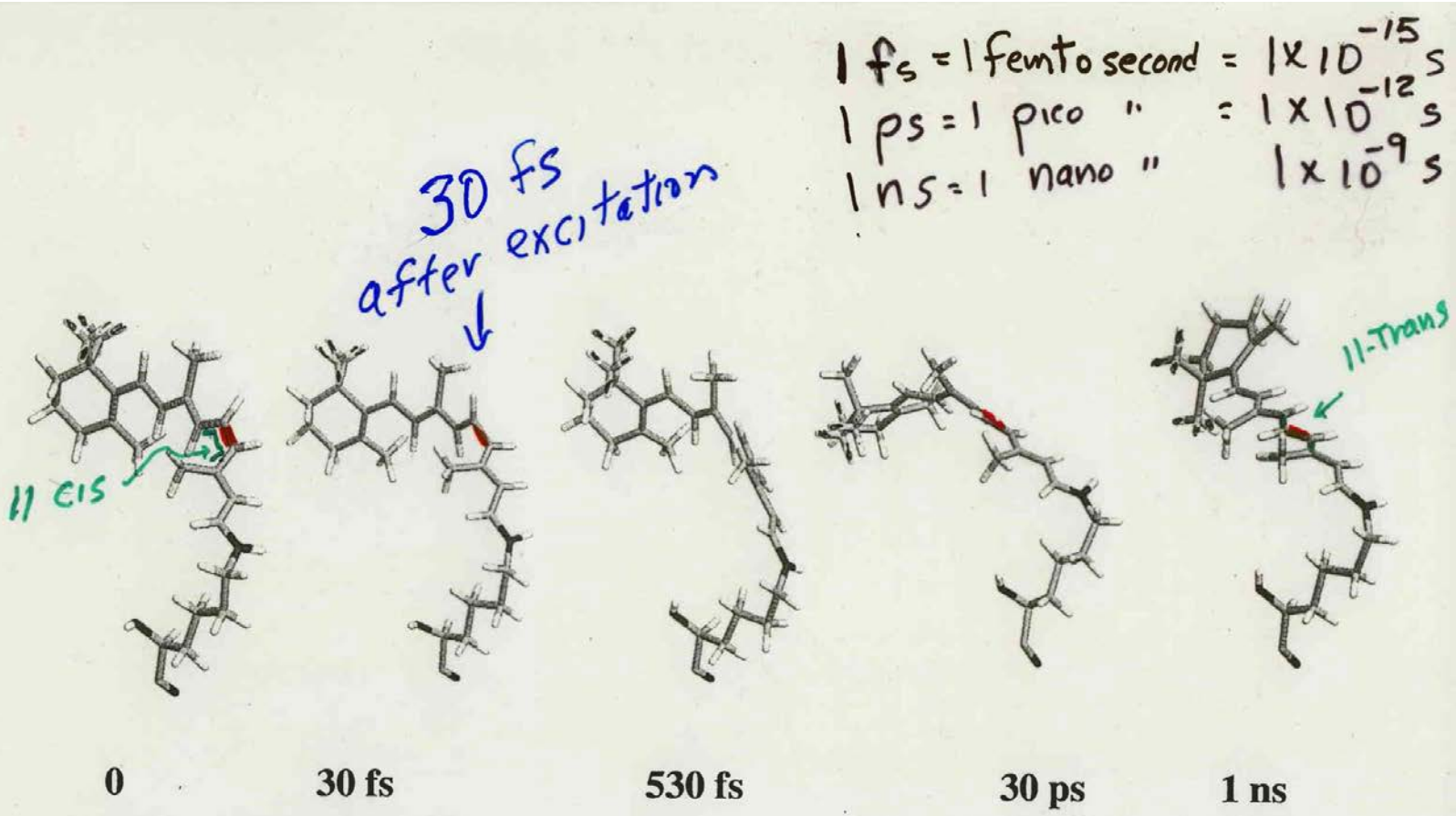
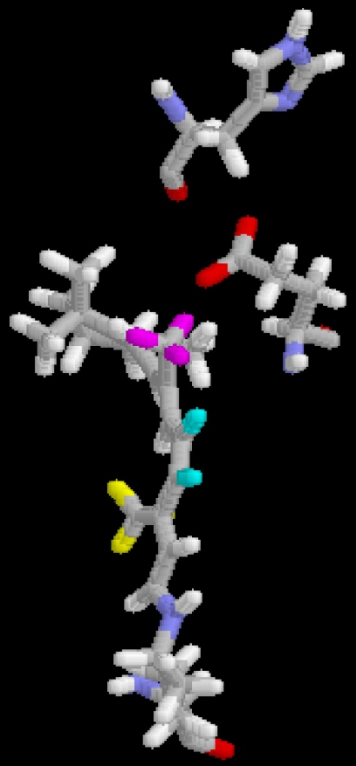
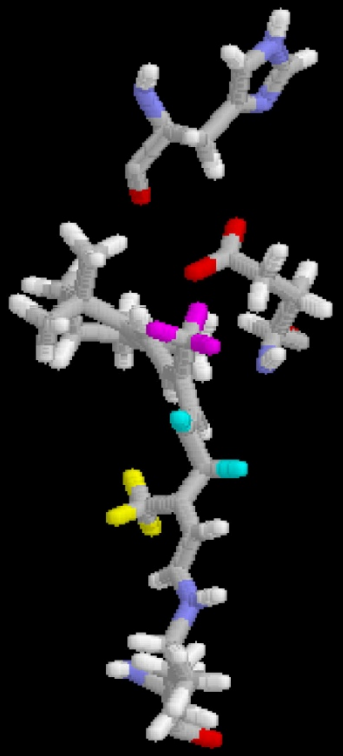


Snapshots of simulated isomerization of retinal Schiff base of rhodopsin (1hzx)

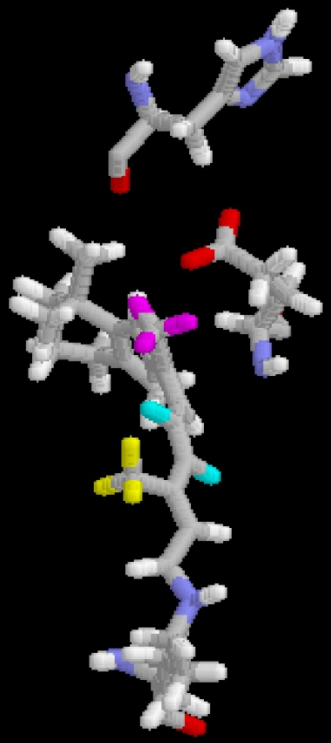
Retinal: 11-cis



excited state



all-trans



Spectral "BUMPS and HUMPS"

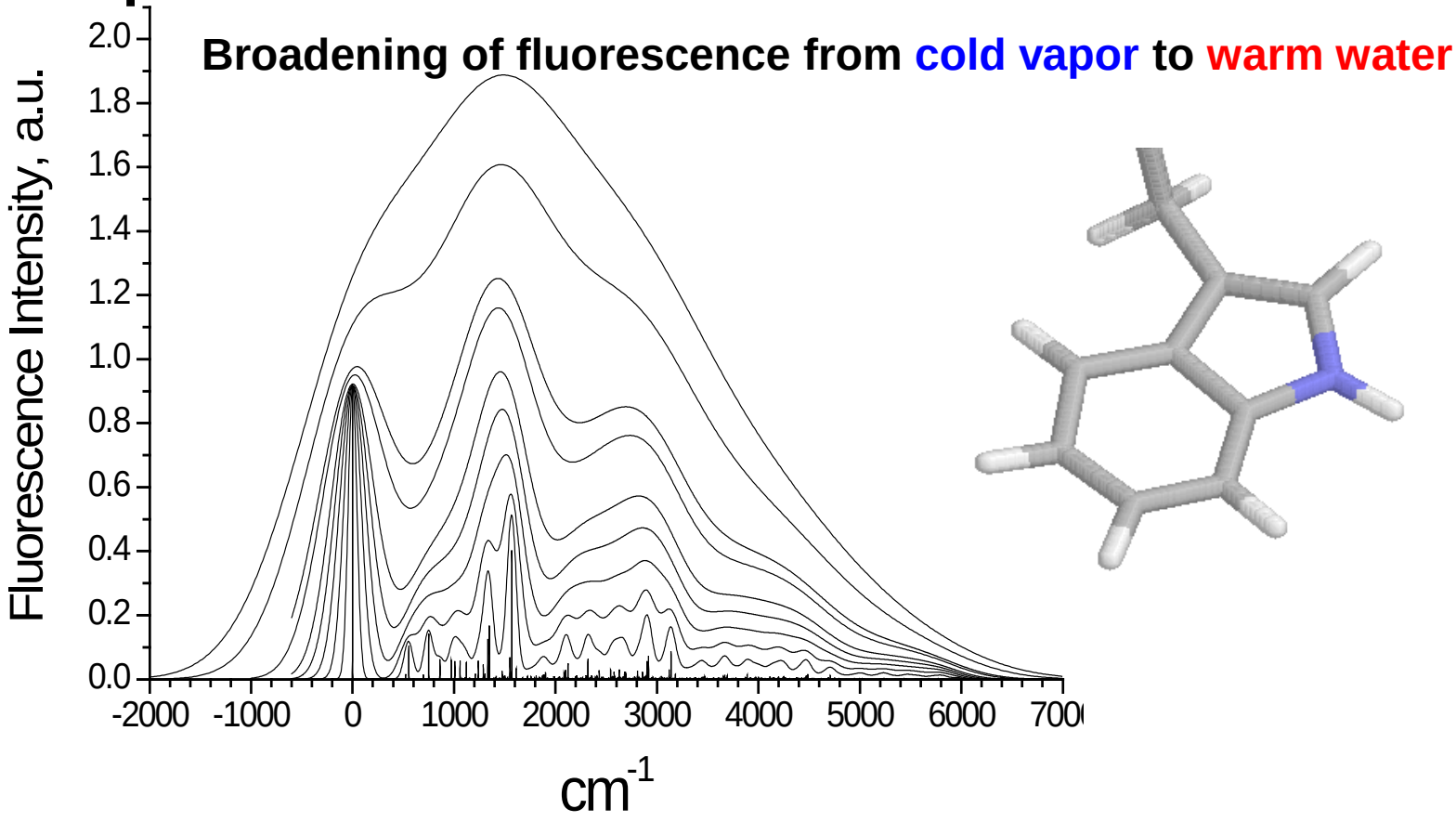
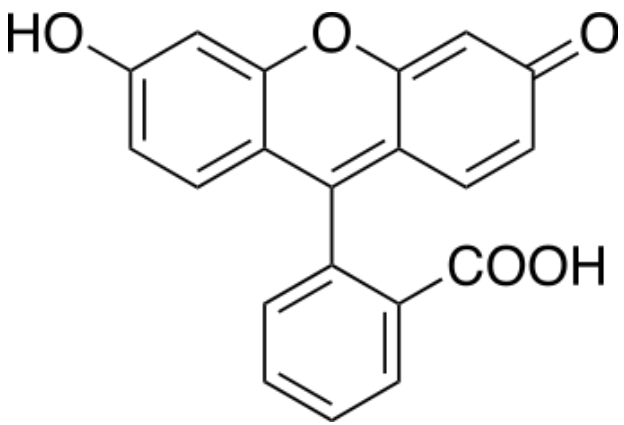


Fig. 21. Broadened calculated spectra. The spectral width of each line of the upwardly displaced spectra is determined by a Gaussian with width of 3, 100, 200, 300, 400, 500, 700, 800, 1200, and 1500 cm^{-1} respectively. The area under each curve is proportional to the line width. (Agrees well with experiment)

Quenching Fluorescein

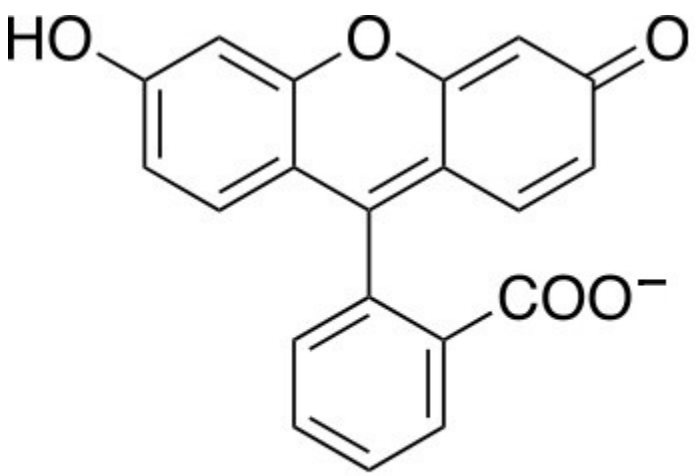
with iodide ion

Fluorescein at different pH



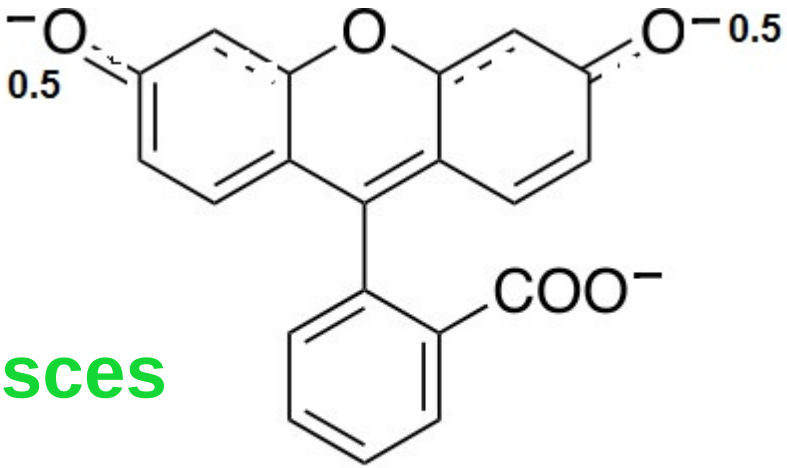
neutral (NOT soluble)
(low pH)

→



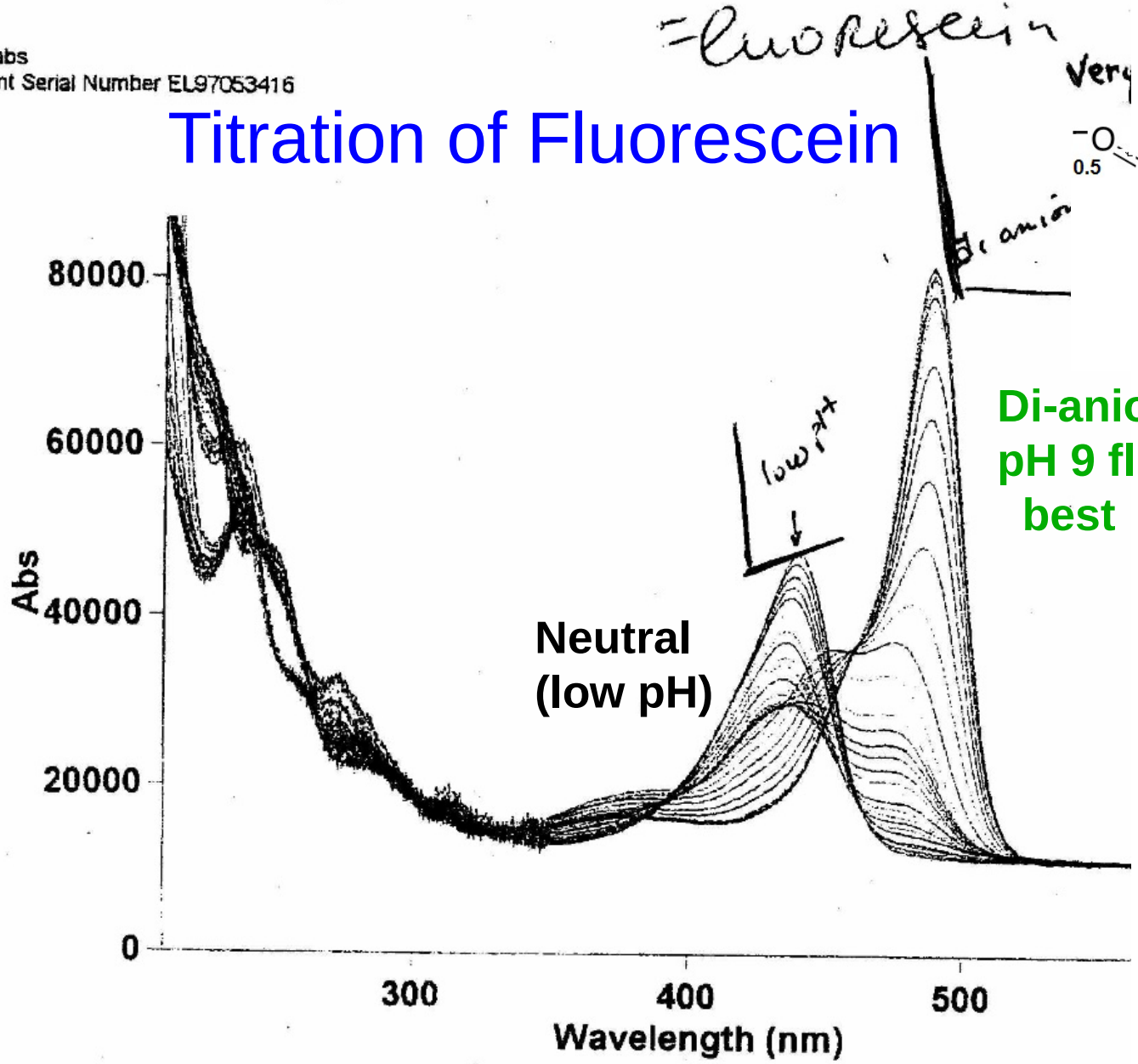
mono-anion (pH 7)

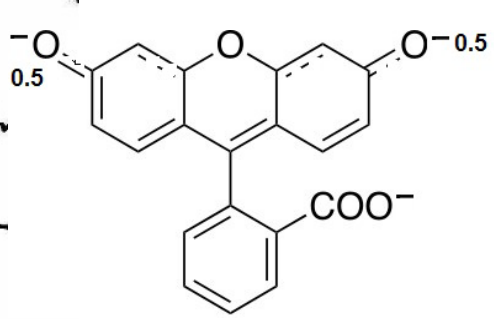
↓



di-anion
pH 9 fluoresces best

Titration of Fluorescein

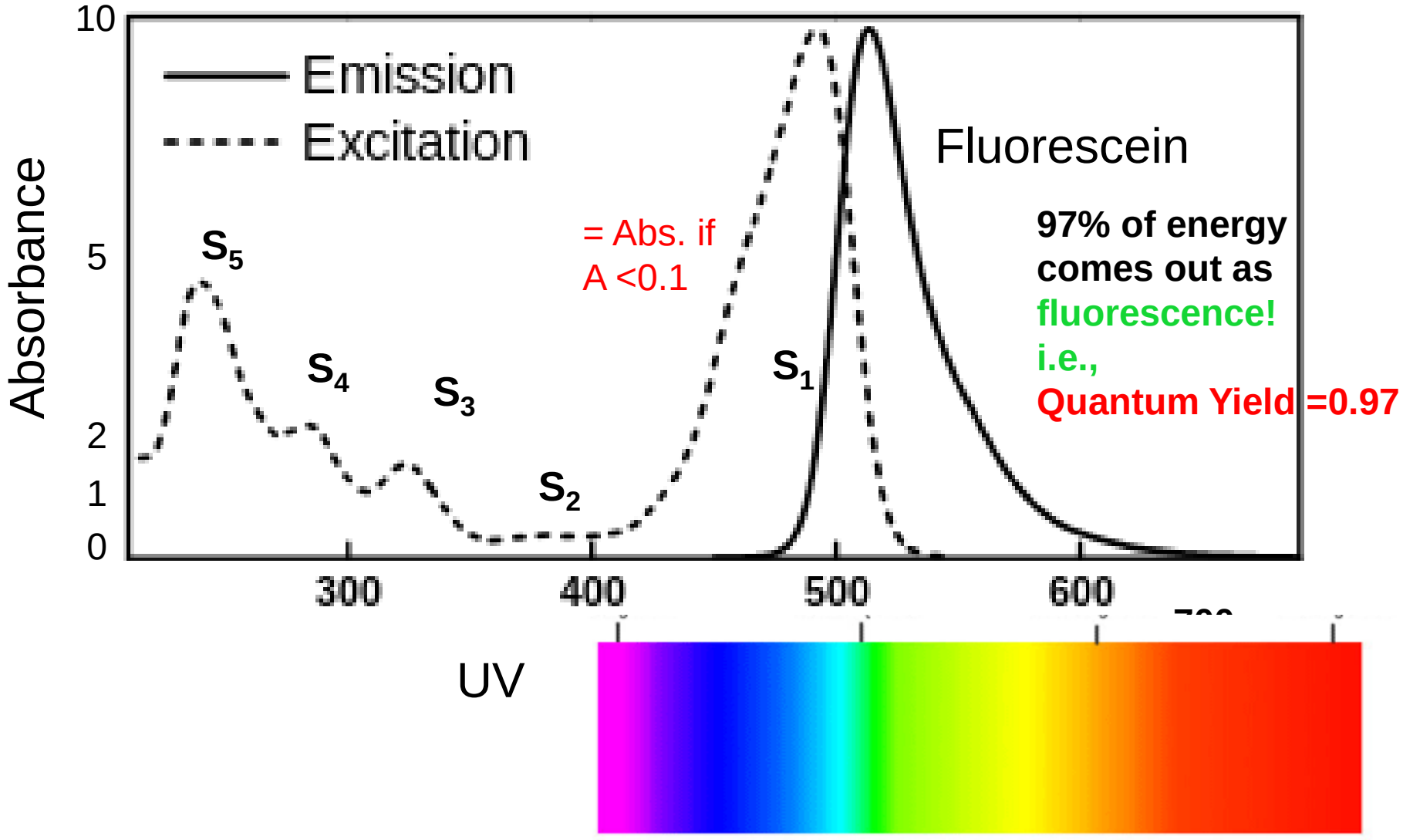


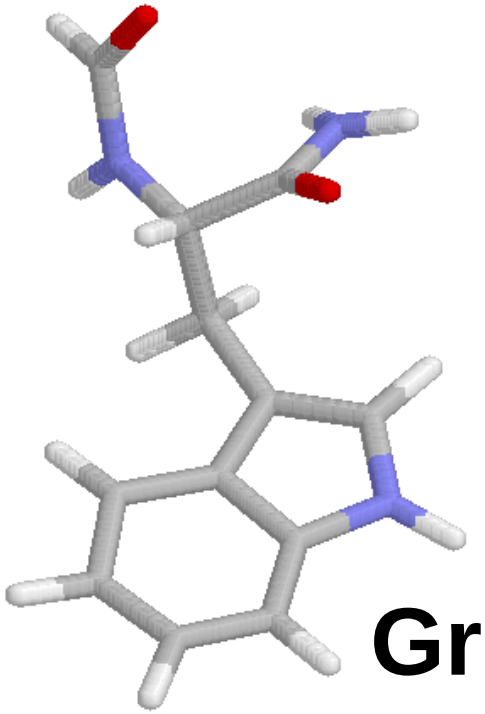


Di-anion
pH 9 fluoresces best

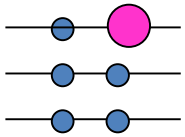
Courtesy Sergey Tetin, Abbott Labs

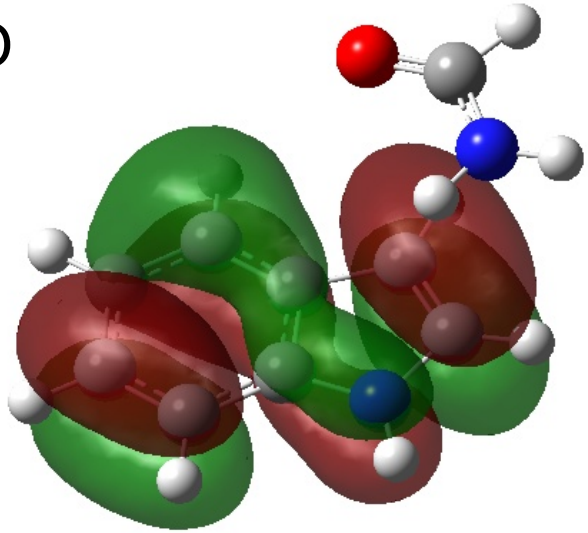
https://en.wikipedia.org/wiki/Fluorescein





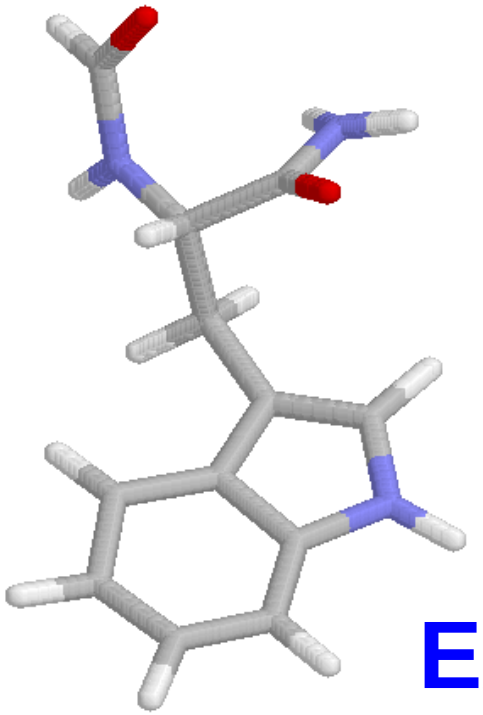
LUMO

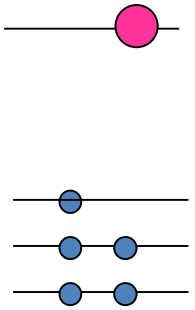


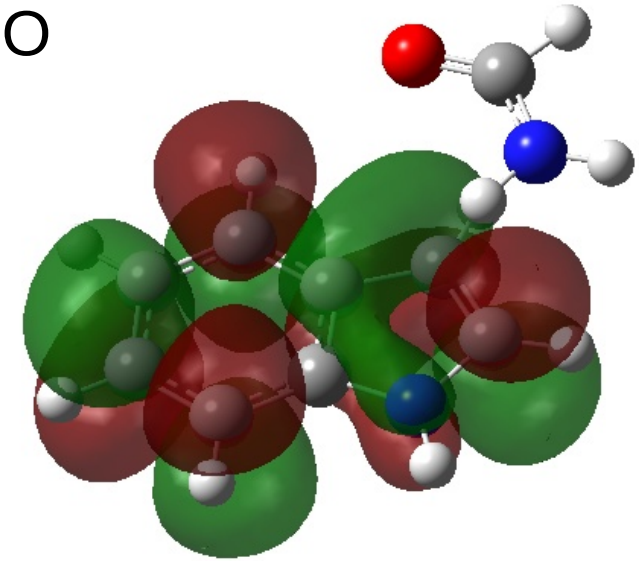


Ground State

Highest Occupied Molecular Orbital



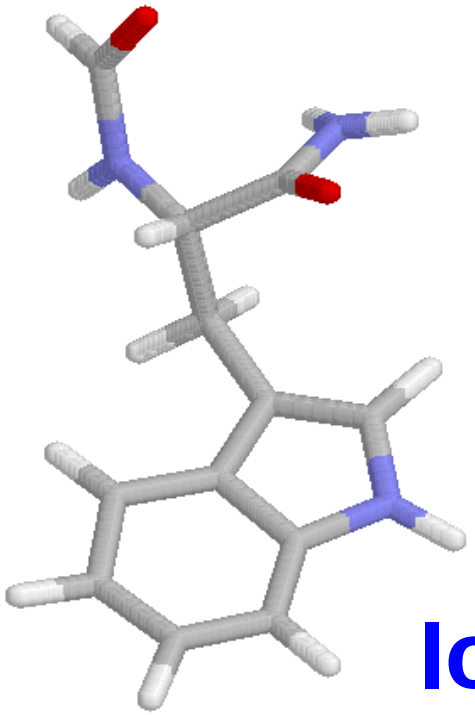


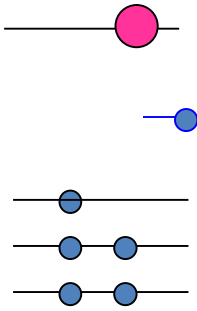


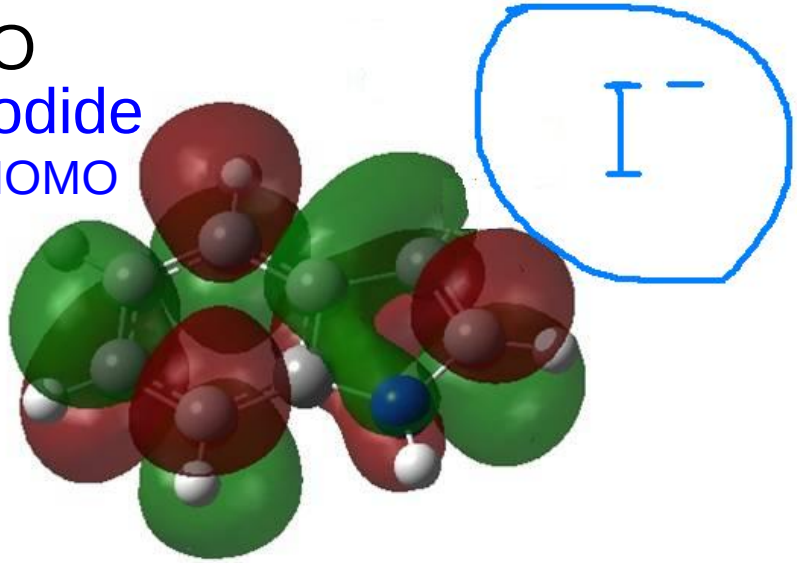
Excited State (fluorescing state)

Lowest Unoccupied Molecular Orbital (electron excited)

Quenching by iodide ion

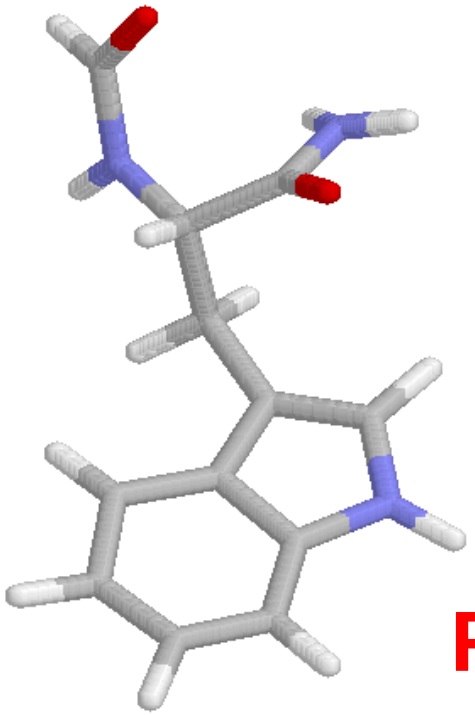


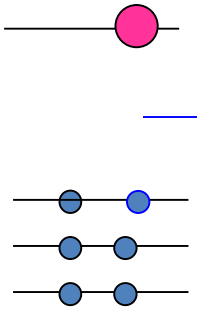


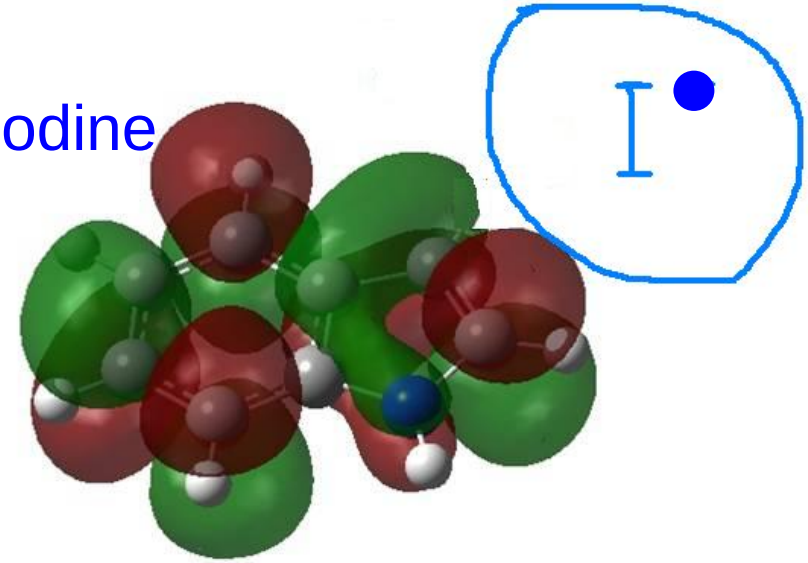


Iodide ion collides;
(has higher HOMO)
will quench fluorescence

Electron transfer from I^- to indole makes a radical pair that cannot fluoresce. (would violate Pauli exclusion)







Radical Pair (can't fluoresce)

Electron transferred from iodide to vacancy in HOMO of ring
i.e., <u>QUENCHING</u>

2.6 Fluorescence Quenching (From Theory Handout)

In fluid solutions, QUENCHING involves a *diffusion controlled*, collisional interaction between the fluorophore and quencher.

What can happen following absorption of a photon?

Absorption: M + hν (photon) → **M*** rate: = ε(λ) [M] x (Intensity of light)

Fluorescence: **M*** → M + hν' rate = k_{rad} [M*]

Internal conversion: **M*** → M + heat rate = k_{ic} [M*]

Intersystem crossing M* → M(triplet) + heat rate = k_{isc} [M*]

Quenching: **M* + Q → M + Q + heat rate = k_q[Q][M*]**

Fluorescence Quantum yield = Φ_f = the rate of fluorescing divided by the total rate of leaving the excited state:

$$\Phi_f = \frac{k_{rad}[M^*]}{(k_{rad} + k_{ic} + k_{isc})[M^*] + k_q[Q][M^*]} = \frac{k_{rad}}{k_{rad} + k_{ic} + k_{isc} + k_q[Q]}$$

For fluorescein, Φ_f = 0.97, τ_{rad} = 4.74 ns so krad= 2.1 x 10^8 k_{ic} + k_{isc} =7 x 10^6 s^{-1}

For quenching by iodide ion, k_q=2 x 10^9 $M^{-1}s^{-1}$.

maximum molar decadic extinction coefficient, ε_{max} = 92,300 $M^{-1}cm^{-1}$ to find the concentration A= εCx

We will measure k_q from the Stern-Volmer Eq.

Rate of collisions = k_q [M*][Q]

Assume every collision quenches. (Diffusion Controlled)

<u>Memorize</u> that this k_q for small molecules is on the order of 10^9 M^{-1} s^{-1} in water.

Depends only on the frictional coefficient (size)

Diffusion coefficient = k_BT/frictional coefficient

We will also make use of the maximum molar decadic extinction coefficient,

ε_{max} = 92,300 $M^{-1}cm^{-1}$ to find the concentration A= εCx

4.2.3 Fluorescence Quenching by Iodide

1. Calculate he quantum yield for each concentration of the iodide, [Q], using numbers from the Theory document and :

Fluorescence Quantum Yield

(a measure of fluorescence ***brightness***)

Quantum Yield = $$\Phi_f = \frac{k_{rad}}{k_{rad} + k_{ic} + k_{isc} + k_q[Q]}$$

Fluorescence Lifetime = 1/(sum of rate constants)

$$= 1/(k_{rad} + k_{ic} + k_{isc} + k_q[Q]) = \tau_f$$

4.2.3 Fluorescence Quenching by Iodide

1. Calculate he quantum yield for each concentration of the iodide, [Q], using numbers from the Theory document and

Compare the measured ratio of the peak height of the quenched cases to the unquenched peak with the calculated values.

$$\Phi_f([Q]) = \frac{k_{rad}}{k_{rad} + k_{ic} + k_{isc} + k_q[Q]} = \text{const x Intensity of fluor} = \text{I}$$

$$\Phi_f \text{ no quencher} = \frac{k_{rad}}{k_{rad} + k_{ic} + k_{isc}} = \text{const x Intensity of fluor} = \text{I}_0$$

$$\frac{I_0}{I} == \frac{(k_{rad} + k_{ic} + k_{isc}) + k_q[Q]}{(k_{rad} + k_{ic} + k_{isc})} = \frac{(k_{rad} + k_{ic} + k_{isc})}{(k_{rad} + k_{ic} + k_{isc})} + \frac{k_q[Q]}{(k_{rad} + k_{ic} + k_{isc})}$$

$= 1 + k_q\tau_0[Q]$, where τ_0 is the unquenched fluorescence lifetime

2., the resulting equation is the widely used Stern-Volmer equation:

where τ_0 is the lifetime in the absence of the quencher, $(k_{rad} + k_{ic} + k_{isc})^{-1}$, I and I_0 are the intensities in the presence and absence of quencher, respectively, and k_q is the diffusion-controlled quenching rate. The product $k_q\tau_0$ is known as the Stern-Volmer constant, K_{sv}.

3. Plot ~~I/I$_0$~~ I_0/I vs. [Q]. Do you get a straight line? What is the Stern-Volmer constant determined from your plot?